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Problems -  
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marina\_lynn8. Key

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this set (58) The  
electrode at which  
oxidation occurs.

anode. One part of a  
voltaic cell in which  
either oxidation or  
reduction occurs. half-  
cell.

Chapter 21

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Tutorial: 1) A voltaic

cell is constructed

with Cd/Cd.  $2+$  at one

half cell and Ag/Ag  $+$

at the other. Both half

cells are Page 2/4

Page 2/6.

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Answers

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G = -nFE<sub>o</sub> CH 17-20 =  
CH15 = CH 21. 2

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Electrochemistry 21.1

= Voltaic Cells (Day

12 = A-4/24, B-4/27)

21.2 = Types of

Batteries (Day 12 =

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Electrolysis (Day 13 =

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A-4/28, B-4/29)

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#43-#46 21.3 =

#20-#24 Unit #14 Test  
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MULTIPLE CHOICE.

Answer Key  
Choose the one  
alternative that best  
completes the  
statement or answers  
the question. 1) The  
gain of electrons by  
an element is  
called\_\_\_\_\_.

- A)oxidation
- B)sublimation
- C)reduction

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D) disproportionation

E) fractionation 2)

           is reduced in  
the following  
reaction:

A.P. Chemistry  
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Section 21.1.

Electrochemistry

represents the

interconversion of  
chemical energy and  
electrical energy.

Electrochemistry invo  
lves redox (reduction-  
oxidation) reactions  
because the electrical  
energy (flow of  
electrons) has at its  
origin the oxidation

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(loss of electrons) and  
reduction (gain of  
electrons) of species.

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questions | Chemistry  
Redox in  
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Electrochemistry is the study of the redox processes by which chemical energy is converted to electrical energy and vice versa.

Electrochemical processes are useful in industry and critically important for biological functioning. In Chapter 19, you read



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that all redox  
reactions involve a  
transfer of  
Answer Key

High-temperature  
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chapters related to  
the fields of

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Electrochemistry

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plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds, biochemical and biological systems, corrosion, energy applications involving fuel cells and solar cells, and nanoscale

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from classical  
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protection systems.

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